Lithium density

Lithium is a solid only about half as dense as water and lithium metal is the least dense metal. A freshly cut chunk of lithium is silvery, but tarnishes in a minute or so in air to give a grey surface. Its chemistry is dominated by its tendency to lose an electron to form Li +. It is the first element within the second period.

Density (g cm -3) Density is the mass of a substance that would fill 1 cm 3 at room temperature. Relative atomic mass The mass of an atom relative to that of carbon-12. This is approximately the sum of the number of protons and neutrons in the nucleus. Where more than one isotope exists, the value given is the abundance weighted average. Isotopes

Lithium has a melting point of 180.54 C, a boiling point of 1342 C, a specific gravity of 0.534 (20 C), and a valence of 1. It is the lightest of the metals, with a density approximately half that of water. Under ordinary conditions, lithium is the least dense of the solid elements.

Lithium has a very low density (0.534 g/cm 3), comparable with pine wood. [13] It is the least dense of all elements that are solids at room temperature; the next lightest solid element (potassium, at 0.862 g/cm 3) is more than 60% denser.

Wikipedia. Lithium is an element number 3 from alkali metals family. Its symbol is Li. Lithium atomic weight is 6.9412 amu. Li stable and long lived isotopes are 6 Li: 7.5%, 7 Li: 92.5%. Lithium electronic configuration is 1s 2 2s 1. Sample compounds that contain Lithium are LiOH, LiCl, Li2CO3, Li3PO4, Li2SO4, LiBr.

Density of Lithium. Density of Lithium is 0.535g/cm 3. Typical densities of various substances are at atmospheric pressure. Density is defined as the mass per unit volume. It is an intensive property, which is mathematically defined as mass divided by volume: r = m/V

Technical data for Lithium. Click any property name to see plots of that property for all the elements. Critical Pressure: Value estimated based on extrapolation. Critical Temperature: Value estimated based on extrapolation. Specific Heat: Value given for solid phase.

Lithium has potential value as a heat-transfer fluid for high power-density nuclear reactors. The lithium-7 isotope, the more common stable isotope, has a low nuclear cross section (that is, it absorbs neutrons very poorly) and thus has potential as a primary coolant for nuclear reactors in which coolant temperatures above about 800 °C (1,500 ...

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